

- Redox reaction between metals produces an electric potential from anode to cathode
- **Rechargeable batteries**- the electron potential can be applied as voltage onto the battery, resulting in a reversal of the reaction initially.

The exact energy can be calculated using the following equation

- $E^{\circ}_{\text{cell}} = E^{\circ}_{\text{cathode}} - E^{\circ}_{\text{anode}}$
- $\Delta G = -nFE^{\circ}_{\text{cell}}$

Examples of Batteries

Non-Rechargeable:

- Zinc-Carbon
- ZnCl
- LiMnO_2
- Alkaline

Rechargeable:

- Lithium Ion
- Lead-Acid
- NiCd
- NiMH

Non

-
-
-
-
-
-
-

Rec

-
-
-