

and it is, therefore, unlikely that after detachment of the first H atom, a second H atom would be eliminated. Little can be said concerning the modes of formation of the pentenes, except that they are related to those of hydrogen.

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Investigation of Some Ce³⁺-Activated Phosphors

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The fluorescence of a number of new Ce³⁺-activated phosphors is described and discussed. Only host lattices with a sublattice consisting of trivalent lanthanide ions are used to avoid charge compensation of the Ce³⁺ ion. Usually the Ce³⁺ emission is in the near-ultraviolet region. Y₃Al₅O₁₂-Ce and SrY₂O₄-Ce, however, show emission in the visible region with a maximum in the green. Conditions for visible Ce³⁺ emission are indicated, viz., large crystal-field splittings (Y₃Al₅O₁₂-Ce) or a large Stokes shift (SrY₂O₄-Ce). In a number of cases we were able to observe all crystal-field components of the excited 5*d* level of Ce³⁺. The cubic crystal-field splitting of the 5*d* level varies strongly with host lattice from 7000 to 14 000 cm⁻¹. The position of the center of the 5*d* levels in oxides is about 30% lower than in the free ion. For some phosphors we observed more than one emission band at room temperature. This is due to fluorescence from higher excited levels. Efficient energy transfer from Ce³⁺ to Cr³⁺ was observed in Y₃Al₅O₁₂.

I. INTRODUCTION

The fluorescence of Ce³⁺-activated compounds is well known.^{1,2} Usually the emission consists of a broad band with two peaks in the long-wavelength ultraviolet region. Since the Ce³⁺ ion has a 4*f*¹ configuration, the ground state consists of a doublet (²F_{5/2} and ²F_{7/2}). The lower excited states are the crystal-field components of the 5*d* configuration. The transition 4*f*→6*s* and charge-transfer transitions are at considerably higher energies³ and are not considered in this paper. The Ce³⁺ emission is due to a 5*d*→4*f* transition. This is an allowed electric dipole transition and therefore the decay time of the fluorescence is very short (<10⁻⁷ sec).^{1,4}

Recently we have also reported on a Ce³⁺-phosphor which emits in the visible, viz., Y₃Al₅O₁₂-Ce.⁴ This peculiar behavior prompted us to study the Ce³⁺ fluorescence in other oxides of the trivalent lanthanides (La³⁺, Gd³⁺, Y³⁺, Sc³⁺). Surprisingly enough, these materials have not widely been investigated, although these host lattices offer the possibility of introducing Ce³⁺ without charge compensation.

The present paper gives the results of this investigation. The complete crystal-field splitting of the 5*d* configuration of the Ce³⁺ ion was found in a number of cases. Requirements for Ce³⁺ emission in the visible region are indicated.

II. EXPERIMENTAL

Samples were prepared by firing intimate mixtures of high-purity oxides (or compounds, which on decomposition yield oxides) at appropriate temperatures in nitrogen. The Ce³⁺ ions were introduced by replacing part of the trivalent lanthanide ions by Ce³⁺ ions. The Ce³⁺ concentration was 1–2 at. %. Products were checked by x-ray analysis. The performance of the optical measurements has been described previously.^{4–6}

III. REFLECTION AND EXCITATION SPECTRA

A. Results

Tables I–III contain the absorption bands of the Ce³⁺ ion in a number of host lattices in the uv region. These host lattices do not absorb in the optical region studied except for the Ga-containing oxides. The data were obtained from excitation and diffuse reflection spectra. In some cases only a continuous absorption was found, so that the absorption bands remain unknown (Table II). Figures 1–3 show the excitation and diffuse reflection spectra of YAl₃B₄O₁₂:Ce, Y₃Al₅O₁₂:Ce, and ScBO₃:Ce. These figures illustrate two types of excitation spectra, viz., those in which the quantum efficiency of the Ce³⁺ fluorescence is roughly independent of the exciting wavelength (YAl₃B₄O₁₂:Ce, Fig. 1) and those in which the quantum efficiency decreases

¹ A. Brill and H. A. Klasens, *Philips Res. Rept.* **7**, 421 (1952).

² J. W. Gilliland and M. S. Hall, *Electrochem. Tech.* **4**, 378 (1966).

³ E. Loh, *Phys. Rev.* **154**, 270 (1967).

⁴ G. Blasse and A. Brill, *Appl. Phys. Letters* **11**, 53 (1967).

⁵ A. Brill and W. L. Wanmaker, *J. Electrochem. Soc.* **111**, 1363 (1964).

⁶ G. Blasse and A. Brill, *J. Chem. Phys.* **46**, 2579 (1967).

TABLE I. Efficiencies for ultraviolet and cathode ray excitation, positions of emission and excitation bands and Stokes shift of the emission of some Ce³⁺ phosphors.

Composition ^a	η (%) ^b	q_{\max} (%) ^c	Emission bands (10 ³ cm ⁻¹) ^d	Excitation bands (10 ³ cm ⁻¹) ^e	Stokes shift (10 ³ cm ⁻¹) ^f
Sc ₂ Si ₂ O ₇ -Ce	1.5	65	25.4; 29.4	29.0; 33.3; ~43.5	3.6
ScBO ₃ -Ce	2	70	26.8; ~30.5	28.0; 31.2; 36.1; 38.5	1.2
YBO ₃ -Ce	2	50	23.8; 25.4	27.4; (~29); 40.8; (~43.5)	2.0
LaBO ₃ -Ce	0.2	35	26.6; 28.4; 31.5	30.8; 36.9; 41.5	2.4
YPO ₄ -Ce	2.5	30	28.3; 30.0	32.8; (34.2); 39.6	2.8
LaPO ₄ -Ce	2.2	40	29.8; 31.5	36.2; 38.8; 41.6	4.7
YAl ₃ B ₃ O ₁₂ -Ce	2	40	27.2; 29.1	31.0; 36.6; 39.2	1.9
Y ₃ Al ₅ O ₁₂ -Ce	3.5	~70	18.2; 27.8; ~29.2	22.0; 29.4; ~37; ~44	3.8
YOCl-Ce	3.5	60	26.3	31.6; 35.8	5.3
LaOCl-Ce	0.4	30	27.8	(~35); 35.8; 39.7	5.2
LaOBr-Ce	0.2	25	22.8	28.4; 34.7	5.6

^a Ce³⁺ concentration is 1-2 at. %.^b Radiant efficiency for cathode ray excitation (20 kV).^c Quantum efficiency for excitation in the lowest excited level.^d Position of the emission peaks.^e Position of the excitation bands; shoulders between brackets.^f Stokes shift of the emission, i.e., difference between the positions of the lowest excited level and the short-wavelength component of the emission from this level.

drastically with decreasing exciting wavelength (Y₃Al₅O₁₂-Ce, Fig. 2 and ScBO₃-Ce, Fig. 3). All the excitation spectra we obtained belong to either the former or the latter type. In this section we discuss the number and position of the absorption bands and a possible explanation for the two types of excitation spectra. A comparison between the excitation spectra of Ce³⁺- and Tb³⁺-activated phosphors is also included.

B. Discussion

The occurrence of more than one Ce³⁺ absorption band in the region between 25 000 and 50 000 cm⁻¹ is due to the crystal-field splitting of the 5d (²D) state. As argued before in the case of Tb³⁺ we can neglect the effect of spin-orbit splitting, since the spin-orbit parameter ζ_{5d} is about 10³ cm⁻¹, whereas the crystal-field splitting amounts to 10⁴ cm⁻¹ or more.⁷ For nearly all compounds mentioned in the tables the symmetry of the lanthanide site is known, so that it is possible to compare the experimentally observed number of 5d levels and the number expected from the site symmetry. Not in all cases a reasonable agreement was found.

Consider, for example, YOCl-Ce and YPO₄-Ce. The respective site symmetries are C_{4v} and D_{2d}, so that 3 and 4 levels are expected. The experimental numbers (Table I) are 2 and 3, respectively. It cannot be excluded, however, that there are levels beyond the spectral limit of our apparatus (~220 m μ) or that lower symmetrical splittings remain hidden in the broad bands observed experimentally. We never observed more bands than expected from the site-symmetry. In Table IV we have tabulated those host lattices for which the number of levels observed experimentally agrees reasonably well with the number expected from site symmetry.

From these data the cubic crystal-field splitting Δ and the center of gravity of the 5d levels can be found, at least approximately. In Y₃Al₅O₁₂-Ce the Ce³⁺ ion occupies a distorted cube, so that the 5d level is split into a lower doublet and a higher triplet by the cubic component of the crystal field. Noncubic components cause a further splitting. Unfortunately we observed only two components of the triplet, which causes some uncertainty in the value of Δ and the position of the center of gravity. In ScBO₃-Ce and Sc₂Si₂O₇-Ce the

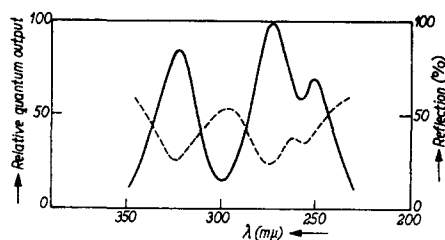


FIG. 1. Diffuse reflection spectrum (broken line) and relative excitation spectrum (solid line) of the Ce³⁺ fluorescence of Y_{0.99}Ce_{0.01}Al₃B₃O₁₂.

⁷ G. Blasse and A. Bril, Philips Res. Repts. **22**, 481 (1967).

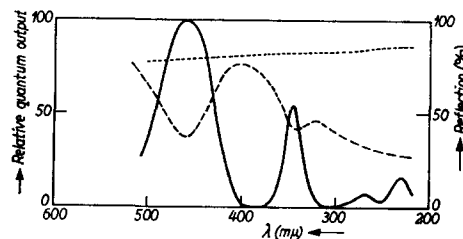


FIG. 2. Diffuse reflection spectrum (broken line) and relative excitation spectrum (solid line) of the Ce³⁺ fluorescence of Y_{2.94}Ce_{0.06}Al₅O₁₂. The dotted line represents the diffuse reflection spectrum of unactivated Y₃Al₅O₁₂.

TABLE II. Position of emission and excitation bands and Stokes shift of the emission of some Ce³⁺ phosphors (inefficient at room temperature).^a

Composition	Emission bands (10 ³ cm ⁻¹)	Excitation bands (10 ³ cm ⁻¹)	Stokes shift (10 ³ cm ⁻¹)
Y ₄ Al ₂ O ₉ -Ce	(~28); 31.2	32.6; ~34; ~41.5	~4.5
SrY ₂ O ₄ -Ce	17.4 ^b	25.2; ?	~8
Y ₂ O ₃ -Ce	19.6; (~24)	?	...
NaYO ₂ -Ce	(~21.3); 22.8	?	...
NaGdO ₂ -Ce	21.3	?	...
LaAlO ₃ -Ce	...	24.2; 31.2; ~40 ^c	...

^a For explanation of columns see Table I.

^b At 77°K.

^c Data from reflection spectrum.

Ce³⁺ ion occupies distorted octahedra, so that the triplet is lower. In YAl₃B₄O₁₂-Ce the Ce³⁺ ion is in a trigonal prism of O²⁻ ions. The ²D level splits into three levels with symmetry species E', A₁', and E'' (31 000, 36 600, and 39 200 cm⁻¹, respectively). This corresponds to a cubic crystal-field splitting of 6500 cm⁻¹ (see Appendix).

Finally Table IV contains analogous data for SrF₂:Ce³⁺ recently given by Loh.³ Although our results are not accurate, Table IV gives some interesting results. The 4f-5d distance is 51 000 cm⁻¹ in the free Ce³⁺ ion. This value decreases to 48 000 cm⁻¹ in fluorides and to roughly 35 000 cm⁻¹ in oxides, a reduction of 6% and 30%, respectively. This fact must be ascribed to the reduction of the interelectronic repulsion parameters by covalency effects (nephelauxetic effect⁸). Within the group of oxidic host lattices there is even a variation of this value. For example, ScBO₃-Ce has the center of the 5d level at ~32 500 cm⁻¹. On the other hand, the value in the case of LaPO₄-Ce must be considerably higher than the 35 000 cm⁻¹ mentioned above (see Table I). In our paper on Tb³⁺-activated phosphors⁷ we have already drawn attention to this fact. The Tb³⁺-activated phosphors show excitation bands due to

4f→5d transitions in the Tb³⁺ ion. These bands are at higher wavenumbers than those of Ce³⁺ and partly beyond the shorter-wavelength limit of our apparatus, so that it was not possible to find the center of gravity in the case of Tb³⁺.

Table IV also shows that the cubic crystal-field splitting Δ varies strongly. For a cube the absolute value of Δ is expected to be 8/9 times the value of Δ for an octahedron with the same distances between central ion and ligand.⁹ With this in mind all values of Δ in Table IV can be compared with each other. The Δ in Y₃Al₅O₁₂-Ce is relatively large, whereas Δ in YAl₃B₄O₁₂-Ce is relatively small. That Δ is strongly influenced by the nature of the next-nearest neighbors is a well-known phenomenon.^{10,11} We further note that Δ in the fluorides has the same order of magnitude as in the oxides. This has also been found for the transition-metal ions.⁸

The value of the noncubic crystal-field splitting varies considerably. For some host lattices it is possible to compare the splitting of the lower cubic crystal-field level of the 5d state of Ce³⁺ and the 4f⁷5d state of Tb³⁺. The agreement is very good (Table V).

The phenomenon that the efficiency of Ce³⁺-activated phosphors is more or less independent of the position of the excitation band in some cases (Fig. 1), but decreases strongly for excitation in bands with higher

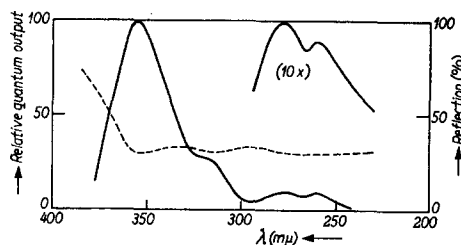
TABLE III. Efficiency for cathode-ray excitation and position of the two lower absorption bands and of the emission band in the visible region for some Ce³⁺-activated garnets.

Composition	η (%)	Position lower absorption bands (difference between brackets) (10 ³ cm ⁻¹)	Emission bands in the visible region (10 ³ cm ⁻¹) ^a
Y _{1.5} Gd _{1.5} Al ₅ O ₁₂ -Ce	... ^b	21.5, 29.6(8.1)	17.4
Y ₃ Al ₅ O ₁₂ -Ce	3.5	22.0, 29.4(7.4)	18.2
Y ₃ Al ₄ GaO ₁₂ -Ce	1.9	22.5, 29.1(6.6)	18.5
Y ₃ Al ₃ Ga ₂ O ₁₂ -Ce	1.7	23.0, 28.8(5.8)	19.2
Y ₃ Al ₂ Ga ₃ O ₁₂ -Ce	1.2	23.3, 28.6(5.3)	19.6(18.3)
Y ₃ Ga ₃ O ₁₂ -Ce	...	23.8, 28.1(4.3)	... ^c

^a Shoulders between parentheses.

^b Emission contains Eu³⁺ contributions due to the starting material Gd₂O₃, which contains a slight amount of europium.

^c No fluorescence.

³ C. K. Jørgensen, *Absorption Spectra and Chemical Bonding* (Pergamon Press, Inc., New York, 1962).

FIG. 3. Diffuse reflection spectrum (broken line) and relative excitation spectrum (solid line) of the Ce³⁺ fluorescence of Sc_{0.99}Ce_{0.01}BO₃.

⁹ J. S. Griffith, *The Theory of Transition-Metal Ions* (Cambridge University Press, Cambridge, 1961).

¹⁰ J. Ferguson, K. Knox, and D. C. Wood, *J. Chem. Phys.* **35**, 2236 (1961); **37**, 193 (1962).

¹¹ G. Blasse, *J. Inorg. Nucl. Chem.* **29**, 1817 (1967).

TABLE IV. The 5*d* levels and cubic crystal-field splitting (Δ) for Ce³⁺ in several host lattices. All values in 10³ cm⁻¹.

Composition	Coordination of Ce ³⁺	5 <i>d</i> levels (observed) ^a	Deduced cubic levels	Cubic crystal-field splitting (Δ)	Spherical 5 <i>d</i>
Y ₃ Al ₅ O ₁₂ -Ce	Distorted cube (<i>D</i> ₂)	22.0 29.4 37 44	25.7 (<i>e_g</i>), ~40 (<i>t_{2g}</i>)	~14	~34.5
YAl ₃ B ₄ O ₁₂ -Ce	Trigonal prism (<i>D</i> _{3h})	31.0 36.6 39.2	...	~6.5 ^o	35.4
ScBO ₃ -Ce	Distorted octahedron (<i>D</i> _{3d})	28.0 31.2 36.1 38.5	~29.5 (<i>t_{2g}</i>), 37.3 (<i>e_g</i>)	~8	~32.5
Sc ₂ Si ₂ O ₇ -Ce	Distorted octahedron (<i>C</i> ₂)	29.0 33.3 43.5	~31 (<i>t_{2g}</i>), ~43.5 (<i>e_g</i>)	~12.5	~35
SrF ₂ -Ce ^b	Distorted cube	33.6 48.8 50.3 53.4	41.2 (<i>e_g</i>), 52.4 (<i>t_{2g}</i>)	11.2	48
Free ion ^b	51

^a Vertical bar indicates separation between lower and higher cubic levels.

^b After Ref. 3.

^o Value of the octahedral crystal-field splitting derived from the position of the levels for Ce³⁺ in prismatic coordination using a point-charge model (see Appendix).

wavenumbers in other cases (Figs. 2 and 3), is probably due to the position of the configuration coordinate curves relative to each other. Let us call the configuration coordinate curves of the ground state and the two lowest excited states *g*, *e*₁, and *e*₂, respectively (*e*₁ is below *e*₂). If the curve *e*₁ crosses the curve *e*₂ within the configuration coordinate curve of *g*, excitation into *e*₁ and *e*₂ may both result in efficient fluorescence from *e*₁. If on extrapolation the curves *e*₁ and *e*₂ should cross outside the curve *g*, however, excitation into *e*₁ gives efficient fluorescence from *e*₁, but excitation into *e*₂ cannot give efficient fluorescence from *e*₁, since either fluorescence from *e*₂ or radiationless transition to *g* results. If the "crossings" of *e*₂ with *e*₁ and *g* approach each other, weak fluorescence from *e*₁ and *e*₂ may be expected upon excitation into *e*₂. From this simple model it follows that a Ce³⁺-activated phosphor with an excitation spectrum in which the efficiency decreases with increasing excitation energy, must show fluorescence from *e*₂ upon excitation into *e*₂ (or higher levels). This was in fact observed (see below).

IV. EMISSION

A. Results

The positions of the emission peaks of our samples are collected in Tables I-III. Figures 4-11 show the spectral energy distribution of the emission of a number

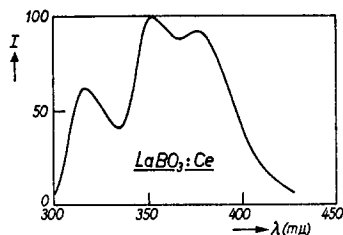


Fig. 4. Spectral energy distribution of the fluorescence of La_{0.99}Ce_{0.01}BO₃ under 254-m μ excitation. In Figs. 5-12 the radiant power per constant wavelength interval in arbitrary units (*I*) is plotted along the ordinate.

of compounds. Usually the emission is in the near-uv region, but not always: Y₃(Al, Ga)₅O₁₂-Ce and SrY₂O₄-Ce (Figs. 5 and 10) show emission in the visible. The half-width value of the emission bands does not depend on their position and amounts to roughly 4000 cm⁻¹ for all materials. In some cases the emission band is split, the difference between the two peaks being some 2000 cm⁻¹ (LaBO₃-Ce, Fig. 4; YBO₃-Ce, Fig. 9; YAl₃B₄O₁₂-Ce and YPO₄-Ce, Table I). In many cases more than one emission band is present (LaBO₃-Ce, Fig. 4; Y₃Al₅O₁₂-Ce, Fig. 5; ScBO₃-Ce, Fig. 6; Sc₂Si₂O₇-Ce, Fig. 7; and Y₄Al₂O₉-Ce, Fig. 11).

B. Discussion

The fluorescence emission of the Ce³⁺ ion originates from a transition from one or more of the 5*d* levels (often only the lowest level) to the ²*F* ground state. Since the ground state of the Ce³⁺ ion is a doublet (²*F*_{5/2} and ²*F*_{7/2}) with a separation of about 2000 cm⁻¹, each emission band is expected to show two peaks as found by us in many cases (compare also Ref. 12). This doublet character of the emission band depends on temperature and Ce³⁺ concentration (self-absorption) and is not always found.¹³

In the case of oxidic host lattices this emission usually peaks in the near uv. As a matter of fact there are two

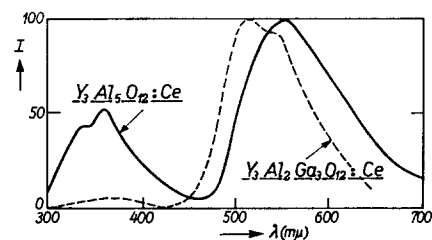


Fig. 5. Spectral energy distribution of the fluorescence of Y_{2.94}Ce_{0.06}Al₅O₁₂ (solid line) and Y_{2.94}Ce_{0.06}Al_{1.2}Ga₃O₁₂ (broken line) under 254-m μ excitation.

¹² F. A. Kröger and J. Bakker, *Physica* **8**, 628 (1941).

¹³ Th. P. J. Botden, *Philips Res. Rept.* **7**, 197 (1952).

TABLE V. Splitting of the lower cubic crystal-field component of the 5*d* level of Ce³⁺ and the 4*f*75*d* level of Tb³⁺ (all values in 10³ cm⁻¹).

Host lattice	Ce ³⁺ ^a	Tb ³⁺ ^b
Y ₃ Al ₅ O ₁₂	7.4	7.5
Y ₃ Ga ₅ O ₁₂	4.3	4.0
YBO ₃	~1.6	1.4
Sc ₂ Si ₂ O ₇	4.3	4.6

^a From Tables I and III.

^b From Ref. 7.

cases in which the emission will be at lower wave numbers.

- If the lowest 5*d* level lies exceptionally low.
- If the Stokes shift of the emission is exceptionally large.

We have found one example of each of these two cases.

In Y₃Al₅O₁₂-Ce the lowest 5*d* level is exceptionally low, viz., at 22 000 cm⁻¹ (see Table I). The Y₃Al₅O₁₂-Ce is the only Ce³⁺-activated phosphor that has not a white, but a yellow body color, whereas the host lattice itself is white. Therefore, the emission lies also at very low wavenumbers. Note in Table I that the Stokes shift of this emission does not have a value deviating from what is normally found. Table IV shows, why the lowest 5*d* level lies so very low. The center of the 5*d* levels of Ce³⁺ in Y₃Al₅O₁₂ has a normal value. In Sec. III we have found that the crystal-field splitting in Y₃Al₅O₁₂ is very large. Moreover, the lower cubic level is the *e_g* level, which is $\frac{3}{5}\Delta$ below the center. In an octahedron the *t_{2g}* level is lower (only $\frac{2}{5}\Delta$ below the center). This means that the lower cubic crystal-field component is relatively low in Y₃Al₅O₁₂ (25 700 cm⁻¹, compared with ~30 000 cm⁻¹ for the lower *t_{2g}* level of ScBO₃-Ce and Sc₂Si₂O₇-Ce). There is yet another effect, viz., the large noncubic splitting of the *e_g* level of Ce³⁺ in Y₃Al₅O₁₂, which brings the lowest level at 22 000 cm⁻¹. Visible Ce³⁺ emission can therefore be expected, if the 5*d* crystal-field splitting and the lower symmetrical splitting are large. Cubic eight coordination is also favorable to obtain visible emission, because the lower cubic level will then be at relatively low energies.

It is obvious that long-wavelength emission will also

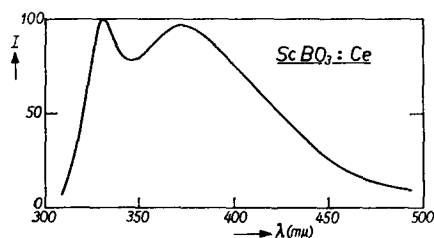


FIG. 6. Spectral energy distribution of the fluorescence of Sc_{0.99}Ce_{0.01}BO₃ under 254-mμ excitation.

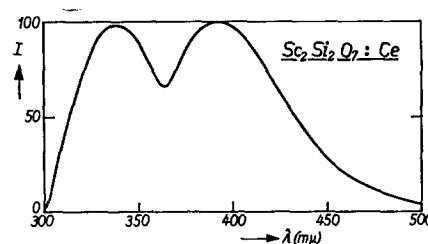


FIG. 7. Spectral energy distribution of the fluorescence of Sc_{1.98}Ce_{0.02}Si₂O₇ under 254-mμ excitation.

result in the case of a very large Stokes shift. This is the case for SrY₂O₄-Ce. The lowest 5*d* level is not exceptionally low (25 200 cm⁻¹), but the Stokes shift is very large (~8000 cm⁻¹, compare Tables I and II), so that visible emission results (Fig. 10). Such a large Stokes shift can only occur if the difference between the equilibrium distance of the excited state and that of the ground state is larger than usual. Elsewhere¹⁴ we have argued that the increase of the equilibrium distance upon excitation will be large if the ions of the host lattice are rather weakly bounded together, i.e., if large and low-charged ions are involved (this implies ions with relatively highly polarization or, as they are sometimes called, "soft" ions¹⁵). Such a situation exists in SrY₂O₄. We also noted weak fluorescence in the visible region in the case of Y₂O₃-Ce, NaYO₂-Ce, and NaGdO₂-Ce (Table II). These are also lattices with large cations. The same holds for YOCl-Ce, LaOCl-Ce, and LaOBr-Ce (Table I), but the oxychlorides emit in the uv. Note, however, that the Stokes shift of these materials is also large.

Another consequence of a large increase of the equilibrium distance upon excitation is a low quenching temperature of the fluorescence. For SrY₂O₄-Ce we found that the light output at 200°K is only 20% of the output at 100°K. Such a marked temperature dependence agrees with our model. In a number of other lattices with large ions (Sr₂LaAlO₅-Ce, SrLaAlO₄-Ce, LaAlO₃-Ce) the Ce³⁺ emission was absent or only very weak, even at liquid-nitrogen temperature.

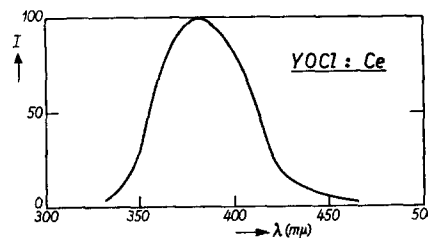


FIG. 8. Spectral energy distribution of the fluorescence of Y_{0.99}Ce_{0.01}OCl under 254-mμ excitation.

¹⁴ G. Blasse and A. Brill, Z. Physik. Chem. (Frankfurt) "The influence of crystal structure on the fluorescence of oxidi niobates and related compounds" (to be published).

¹⁵ Compare C. K. Jørgensen, Coord. Chem. Rev. 1, 164 (1966).

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