HANDBOOK OF

PHARMACEUTICAL

EXCIPIENTS

THIRD EDITION



ARTHUR H. KIBBE

OUTDATED

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Handbook of PHARMACEUTICAL EXCIPIENTS

Third Edition

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Citric Acid Monohydrate

1. Nonproprietary Names

BP: Citric acid monohydrate

JP: Citric acid

PhEur: Acidum citricum monohydricum

USP: Citric acid

2. Synonyms

2-Hydroxypropane-1,2,3-tricarboxylic acid monohydrate.

3. Chemical Name and CAS Registry Number

2-Hydroxy-1,2,3-propanetricarboxylic acid monohydrate [5949-29-1]

4. Empirical Formula

Molecular Weight

C₆H₈O₇·H₂O

210.14

5. Structural Formula

6. Functional Category

Acidifying agent; antioxidant; buffering agent; chelating agent; flavor enhancer.

7. Applications in Pharmaceutical Formulation or Technology

Citric acid, as either the monohydrate or anhydrous material, is widely used in pharmaceutical formulations and food products primarily to adjust the pH of solutions. Citric acid monohydrate is used in the preparation of effervescent granules while anhydrous citric acid is widely used in the preparation of effervescent tablets.⁽¹⁾

In food products, citric acid is used as a flavor enhancer, for its tart, acid taste. Citric acid monohydrate is also used as a sequestering agent and antioxidant synergist. It is a component of anticoagulant citrate solutions. Therapeutically, preparations containing citric acid have been used to dissolve renal calculi.

Use	Concentration (%)	
Buffer solutions		
Flavor enhancer for liquid formulations	0.3-2.0	
Sequestering agent	0.3-2 0	

8. Description

Citric acid monohydrate occurs as colorless or translucent crystals, or as a white crystalline, efflorescent powder. It is odorless and has a strong acidic taste. Crystal structure is orthorhombic.

9. Pharmacopeial Specifications

Test	JP	PhEur	USP
Identification	+	+	+
Clarity and color of solution	2	+	2
Water			
(hydrous form)	_	7.5-9.0%	$\leq 8.8\%$
(anhydrous form)	_	≤ 1.0%	≤ 0.5%
Bacterial endotoxins	-	+	Kara and a second
Residue on ignition	_		≤ 0.05
Sulfated ash	≤ 0.1%		$\leq 0.1\%$
Barium	_	+	()
Calcium	+	≤ 200 ppm	
Aluminum	_	+	3
Oxalate	+	≤ 350 ppm	+
Sulfate	$\leq 0.048\%$	≤ 150 ppm	+
Arsenic	≤ 1 ppm	-	≤ 3 ppm
Heavy metals	≤ 10 ppm	≤ 10 ppm	$\leq 0.001\%$
Iron		≤ 50 ppm	STA
Chloride	_	≤ 50 ppm	_
Readily carbonizable substances	+	+	+
Polycyclic aromatic hydrocarbon	+	1,000	-
Organic volatile impurities	-	 -	+
Assay (anhydrous basis)	≥ 99.5%	99.5-101%	99.5-100.5%

10. Typical Properties

Acidity/alkalinity:

pH = 2.2 (1% w/v aqueous solution)

Dissociation constants:

pK_{a1}: 3.128 at 25°C;

pK_{a2}: 4.761 at 25°C;

pKa3: 6.396 at 25°C.

Density: 1.542 g/cm³

Heat of combustion:

-1972 kJ/mol (-471.4 kcal/mol)

Heat of solution:

-16.3 kJ/mol (-3.9 kcal/mol) at 25°C

Hygroscopicity: at relative humidities less than about 65% citric acid monohydrate effloresces at 25°C, the anhydrous acid being formed at relative humidities less than about 40%. At relative humidities between about 65-75%, citric acid monohydrate absorbs insignificant amounts of moisture but under more humid conditions substantial amounts of water are absorbed.

Melting point: ~ 100°C (softens at 75°C)

Particle size distribution: various grades of citric acid monohydrate with different particle sizes are commercially available.

Solubility: soluble 1 in 1.5 parts of ethanol (95%) and 1 in less than 1 part of water; sparingly soluble in ether.

Viscosity (dynamic): 6.5 mPa s (6.5 cP) for a 50% w/v aqueous solution at 25°C

See also Section 18.



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